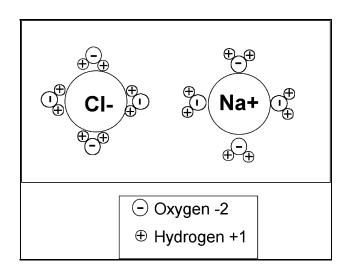
$NaCl + H_2O - Na^+ + Cl^- + H_2O$



- D. Measuring concentrations of solutes in aqueous solutions
 - 1. concentration measurement of the quantity of solute in a given quantity of solvent (or solution)
 - a. Mass Percent = (mass solute / total mass solution) * 100%

E.g. if 5 g of NaCl is dissolved in 95 g of water, what is the mass percent of sodium chloride in the solution?

conversion factors for mass: 1 gram = 1000 mg, 1 kg = 1000 g, 1 gram = 1,000,000 micrograms

- (1) percent = "parts per hundred" (%)
- b. Parts per Thousand (o/oo) = grams of solute / liter of water

Determine the concentration in ppt for a solution of 200 gram dissolved in 2 liters of water?

Determine the concentration in ppt for a solution of 2000 mg dissolved in 1 liter of water?

c. Parts per Million = milligrams of solute / liter of water

Determine the concentration in ppm for a solution 20 mg of salt per liter of water?

What about 20 kg of salt per liter of water?